Unit 11 - Week 9

Assignment 9

Problem 1: Consider the following statements and select the correct option.

Statement A: The boiling point of AgCl is greater than ZnCl₂.
Statement B: The boiling point of AgCl is greater than CuCl²⁻.
Statement C: The boiling point of CuCl²⁻ is greater than ZnCl₂.

Statement D: The boiling point of AgCl is greater than CuCl²⁻.

Which statement is correct?

Problem 2: Consider the following statements and select the correct option.

Statement A: The freezing point of AgCl is greater than ZnCl₂.
Statement B: The freezing point of AgCl is greater than CuCl²⁻.
Statement C: The freezing point of CuCl²⁻ is greater than ZnCl₂.

Statement D: The freezing point of AgCl is greater than CuCl²⁻.

Which statement is correct?

Problem 3: Consider the following statements and select the correct option.

Statement A: The density of AgCl is greater than ZnCl₂.
Statement B: The density of AgCl is greater than CuCl²⁻.
Statement C: The density of CuCl²⁻ is greater than ZnCl₂.

Statement D: The density of AgCl is greater than CuCl²⁻.

Which statement is correct?

Problem 4: Consider the following statements and select the correct option.

Statement A: The solubility of AgCl is greater than ZnCl₂.
Statement B: The solubility of AgCl is greater than CuCl²⁻.
Statement C: The solubility of CuCl²⁻ is greater than ZnCl₂.

Statement D: The solubility of AgCl is greater than CuCl²⁻.

Which statement is correct?

Problem 5: Consider the following statements and select the correct option.

Statement A: The equilibrium constant for the reaction Ag⁺ + Cl⁻ ⇌ AgCl is greater than Zn⁺ + Cl⁻ ⇌ ZnCl₂.
Statement B: The equilibrium constant for the reaction Ag⁺ + Cl⁻ ⇌ AgCl is greater than Cu⁺ + Cl⁻ ⇌ CuCl²⁻.
Statement C: The equilibrium constant for the reaction Cu⁺ + Cl⁻ ⇌ CuCl²⁻ is greater than Zn⁺ + Cl⁻ ⇌ ZnCl₂.

Statement D: The equilibrium constant for the reaction Cu⁺ + Cl⁻ ⇌ CuCl²⁻ is greater than Ag⁺ + Cl⁻ ⇌ AgCl.

Which statement is correct?

Problem 6: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 7: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 8: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 9: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 10: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 11: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?

Problem 12: Consider the following statements and select the correct option.

Statement A: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).
Statement B: The net ionic equation for the following reaction is Cu⁺ + Cl⁻ ⇌ CuCl₂(s).
Statement C: The net ionic equation for the following reaction is Zn⁺ + Cl⁻ ⇌ ZnCl₂(s).

Statement D: The net ionic equation for the following reaction is Ag⁺ + Cl⁻ ⇌ AgCl(s).

Which statement is correct?