

Wendy
5 Dec 2012
10:00-11:30

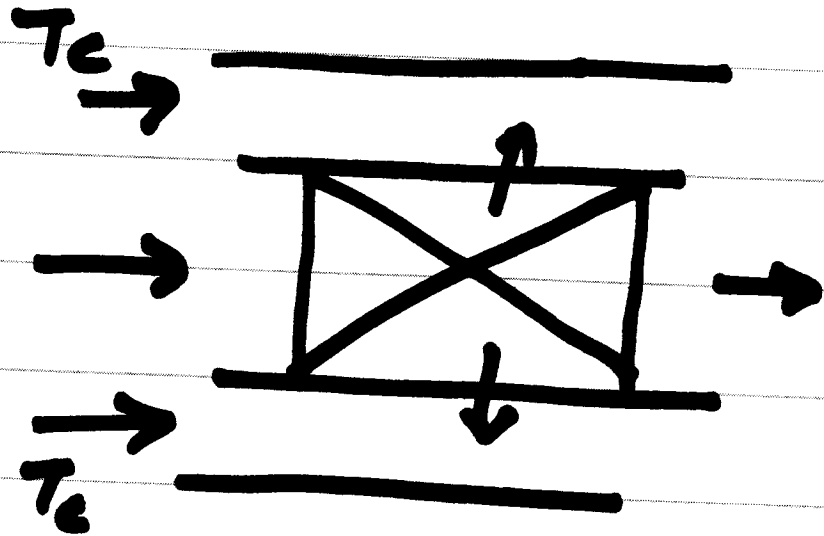
Prof. Shankar

Lec. 34 05/12/2012

Advanced Reaction Engineering

Further Considerations in

Energy Balance



$$\begin{aligned}
 k \text{ at } 513 \text{ K} &= 1800/\text{hr} \\
 E &= 20,000 \text{ cal/mol} \\
 \Delta H &= -152 \text{ kcal/mol} \\
 h &= 20 \text{ kcal/m}^2 \text{ hr} \\
 r &= k C_H \\
 T_c &= 250
 \end{aligned}$$

Energy Balance

$$v C_p \frac{dT}{dv} = r_1 (-\Delta H^*) + \frac{4h}{D} (T_c - T)$$

v - volumetric flow m^3/hr

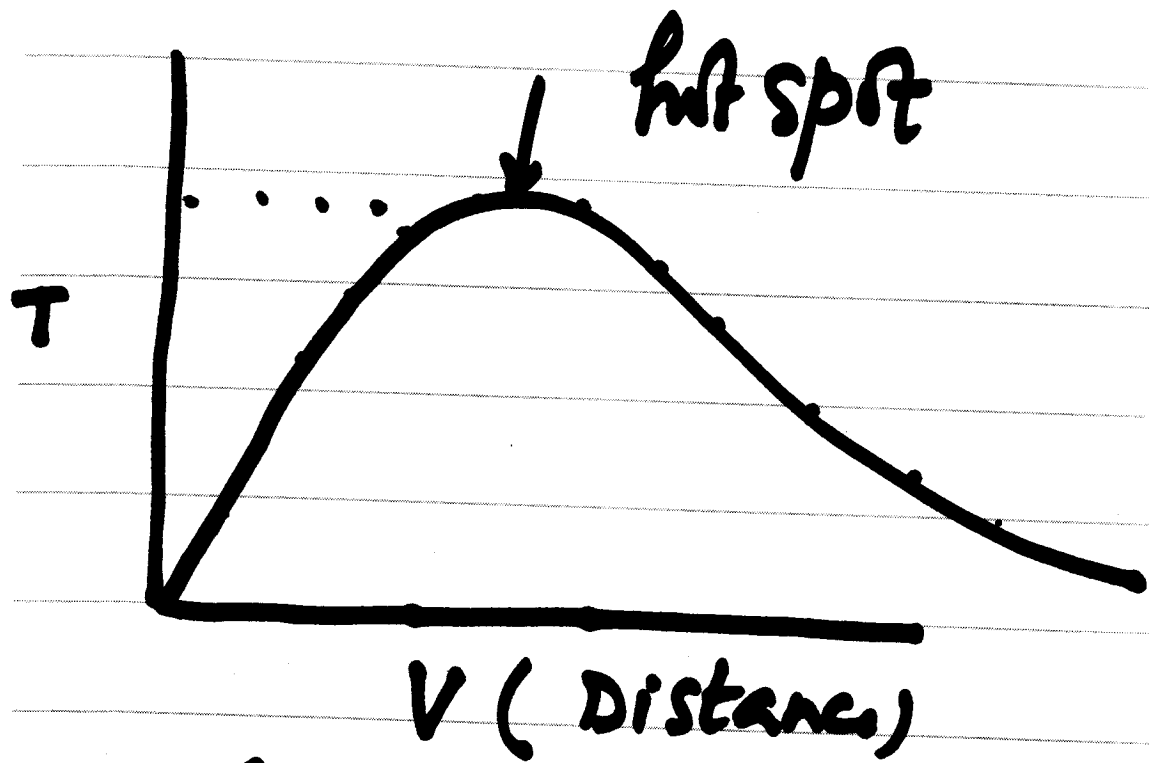
C_p - vol. sp. heat $\text{kcal}/\text{L}\cdot\text{hr}\cdot\text{C}$

r_1 - rate function $\text{mol}/\text{L}\cdot\text{hr}$

h - heat transfer coefficient $\text{kcal}/\text{m}^2\cdot\text{hr}\cdot\text{C}$

D - pipe inside diameter m

T_c - Coolant Temp C



Wall Cooled
Reactor



$\cancel{v C_p} \frac{dT}{dV}$
A hot spot

$Q_r (-\Delta H^*)$
heat gen

$+ \frac{4h(T_c - T)}{D}$
(heat removal)

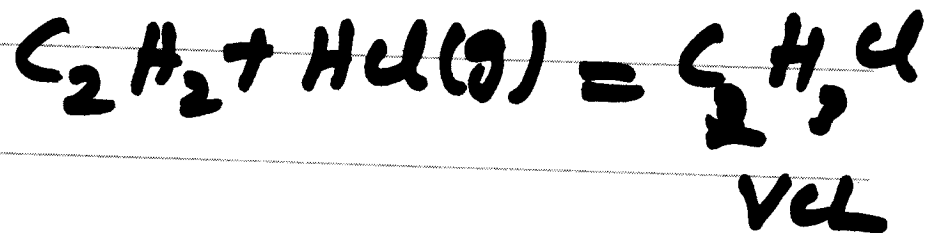
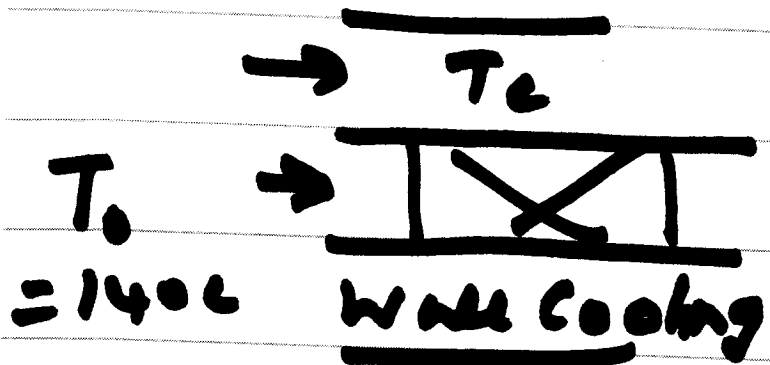
At hot spot we have

$$0 = R_1(-\Delta H^*) + \frac{4h}{D}(T_c - T).$$

$$\underbrace{\frac{4h}{D}(T - T_c)}_{\text{heat removal}} = \underbrace{R_1(-\Delta H^*)}_{\text{heat generation}} \quad \text{at hot spot.}$$

$$P_t = 1 \text{ atm}$$

$$y_{A0} = 0.1 \text{ (HCl)}$$



d_p (m)	T (K or $^\circ\text{C}$)	X_{expt}	y_{A0}
0.05	240	0.43	0.1
0.075	240	0.79	0.1
0.125	300	0.61	0.2

At the hot spot

$$q_1 (-\Delta H^*) = \frac{4h}{D} (T - T_c)$$

$$q_1 = h C_H = h C_{\infty} (1-x)$$

$$h C_{\infty} (1-x) (-\Delta H^*) = \frac{4h}{D} (T - T_c)$$

$$1-x = \frac{4h (T - T_c)}{D h C_{\infty} (-\Delta H^*)}$$

D	T	X
0.05	240	0.48

$$G_A = \frac{P(0.1)}{(RT)} \uparrow$$

$$= (2.34) 10^{-3} \frac{\text{mol}}{\text{L}}$$

$$1-X = \frac{(4)(20)(240-25)}{(0.05)(1800)(2.34) 10^{-3}(152000)}$$

$$1-X = 0.537 \Rightarrow X = 0.46$$

$$1-x = \frac{4h(T-T_c)}{\Delta k C_{A0}(-\Delta H_1^*)}$$

$$= \frac{(4)(20)(240-25)}{\dots}$$

$$C_{A0} = 0.0047 \frac{\text{mol}}{\text{L}}$$

$$(0.075)(1800)(0.0047)(152000)$$

$$1-x = 0.178 \Rightarrow x = 0.82$$

$$1-x = \frac{4h(T-T_c)}{DkC_{A0}(-\Delta H^*)}$$

$$= \frac{(4)(20)(300-25)}{(0.025)(7.76)(1800)(1.05)10^{-3}}$$

$$1-x = 0.394$$

$$x = 0.60$$

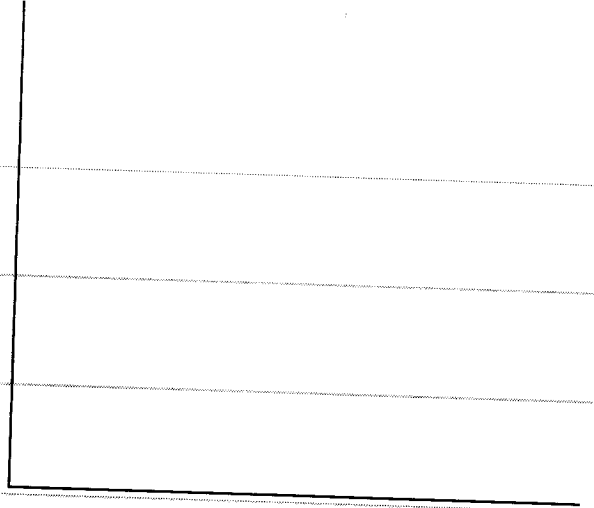
$$k(300) = 7.76$$

$$k(240)$$

$$C_{A0}(300) = (1.05)10^{-3} \frac{\text{mol}}{\text{L}}$$

RA s/pa

M	X_{expt}	X_{Model}	T
.05	0.43	✓ 0.46	240
0.025	0.79 ✓	0.82	240
0.025	0.61	0.60	300



$$k = 1800/\text{hr}$$

Endothermic Reaction.

Processing at $20 \text{ m}^3/\text{hr}$.

Heat of rxn: $20 \text{ kcal}/\text{m}^3$.

Heat ~~balance~~ _{supplies}: $400 \text{ kcal}/\text{hr}$

$$k = 1800/\text{hr}$$

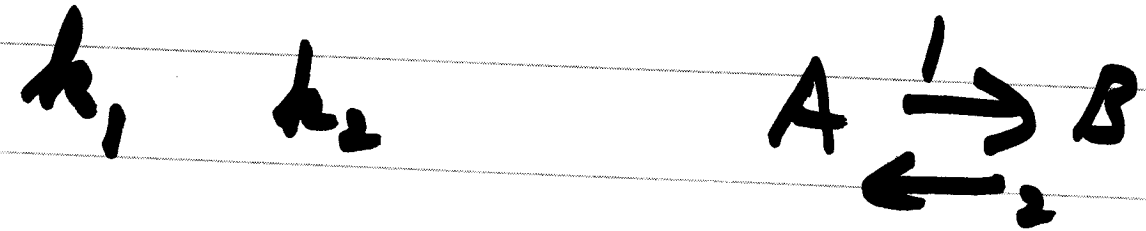
Heat transfer design

$$20 \text{ heat/hr}$$

Heat:

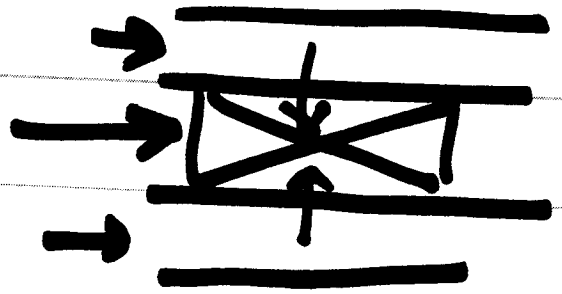
Rxn proceeds at the rate at which heat is supplied.

Endothermic Reaction
which is instantaneous



Rxn is essentially at equilibrium

$(X)_{eq} = \frac{K}{(K+1)} \quad K(T)$



instantaneous

$$A \frac{dx}{dt}$$

$$\begin{aligned} n C_p \frac{dT}{dv} &= r_1 (-\Delta H_1^*) \\ &\quad + r_2 (-\Delta H_2^*) + \frac{4h}{D} (T_2 - T) \\ &= \underline{(r_1 - r_2)} (-\Delta H_1^*) + \frac{4h}{D} (T_2 - T) \end{aligned}$$

$$\frac{dF_A}{dv} = (r_2 - r_1)$$

$$F_{A0} \frac{dx}{dv} = (r_1 - r_2)$$

$$\frac{dx}{dv} = \frac{dx}{dT} \cdot \frac{dT}{dv}$$

$$x = \frac{K}{K+1}$$

Reactions at
instantaneous do
are at equilibrium.

$$\frac{dx}{dT} = \left(\frac{1}{K+1} - \frac{K}{(K+1)^2} \right) \frac{dK}{dT} \cdot \frac{1}{(K+1)^2} \left(\frac{\Delta H}{RT^2} \right)$$

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Vant Hoff's Eqn

$$\frac{d \ln K}{dT} = \frac{\Delta H}{RT^2}$$

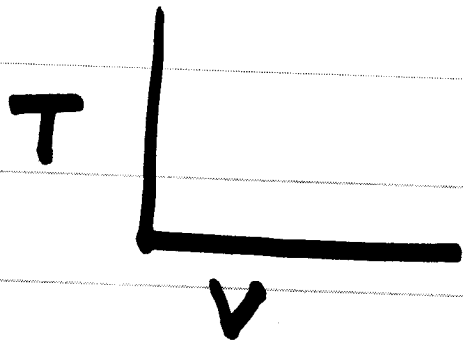
$$\frac{dK}{dT} = K \cdot \left(\frac{\Delta H}{RT^2} \right)$$

$$\underline{\underline{(K/K_{+1}) = X}}$$

$$v C_p \frac{dT}{dv} = (-\Delta H^\ddagger) F_A \frac{dx}{dv}$$

$$+ \frac{4h}{D} (T_c - T)$$

$$v C_p \frac{dT}{dv} = (-\Delta H^\ddagger) F_A \left(\frac{dx}{dT} \right) \frac{dT}{dv} + \frac{4h}{D} (T_c - T)$$



$$v C_p \frac{dT}{dv} = (-\Delta H_i^*) F_{A_0} \frac{(\Delta H_i^*)}{RT^2 (K_2 + 1)^2} \frac{dT}{dv}$$

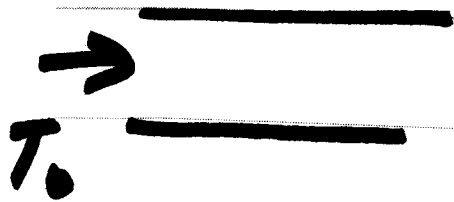
$$+ \frac{4h}{D} (T_c - T).$$

$$\frac{dT}{dv} \left\{ v C_p + \frac{(\Delta H_i^*)^2 F_{A_0}}{RT^2 (K_2 + 1)^2} \right\} = \left[\frac{4h}{D} (T_c - T) \right]$$

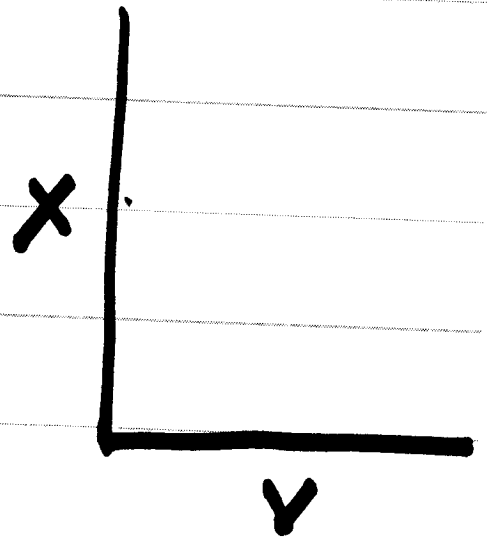
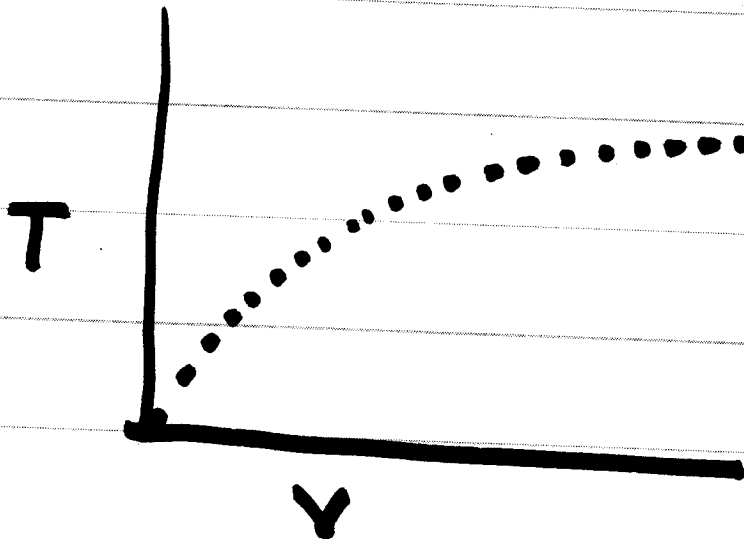
$$= f(T).$$

$$\frac{dT}{dy} = \frac{4h(T_c - T)}{D}$$

$$\left[uG + (\Delta H_v)^2 F_{10} / RT^2 (K+1) \right]$$



$$X = \left(\frac{K}{K+1} \right)$$



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① Instantaneous Rxn

$$X = \frac{K \cdot}{K+1}$$

② Reaction $k = 1800/h$ ←

Endothermic

Heat loss Ret. Control: 10/h